

**Stoichiometry Stumper 1 Answer**

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**Stoichiometry Stumper 1 Answer**  
=4.423g ethyl benzene stoichiometry stumper #1 So, Suzanne did not have enough benzene to create the amount of ethyl benzene she claimed to have made. But she did have enough benzene to create the same amount of di-chloro benzene that killed her husband.

**stoichiometry stumper #1 by Amy Hinds - Prezi**  
Stoichiometry Stumper #1 You are a forensic scientist. You are investigating a murder involving poison. The victim was poisoned with a compound called di-chloro benzene whose formula is C6H4Cl2.

**Solved: Stoichiometry Stumper #1 You Are A Forensic Scient ...**  
Stoichiometry Stumper #1 You are a forensic scientist. You are investigating a murder involving poison. The victim was poisoned with a compound called di-chloro benzene whose formula is C 6 H 4 Cl 2. Autopsy results show that the victim's body contained about 31 g of the poison, but the actual amount could have been slightly

**Stoichiometry Stumper #1 - Tumwater School District**  
Stoichiometry Stumper 1 Answer Forensic Author: electionsdev.calmatters.org-2020-10-22T00:00:00+00:01 Subject: Stoichiometry Stumper 1 Answer Forensic Keywords: stoichiometry, stumper, 1, answer, forensic Created Date: 10/22/2020 5:50:48 PM

**Stoichiometry Stumper 1 Answer Forensic**  
Stoichiometry Stumper Stoichiometry Stumper #4 Step 1 NASCAR pit crew member is suppose to tell the driver how much fuel he has left and if he needs to make another pit stop before he finishes the race. Step 2 The car has 25 gallons or 3.5 kg of fuel. The reaction uses 300 g of

**Stoichiometry Stumper by shia hines - Prezi**  
\$1.25 Monday, February 17, 2014 Procedure: Is she lying? 1. Convert 15 grams into moles by dividing by its mass, 78.114 g C6H6. 2. Multiply the moles by the mass of C6H5CH3. 3. You will get an answer of 18 g. Math Three forensic scientists are investigating a murder involving

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**Stoichiometry Stumper 1 Answer - webmail.bajanusa.com**  
The Answer Our company does not have enough funds to produce the calcium nitrate needed to make the required amount of the final drug. Our company will have to sell the rights to the drug to a more well-funded company because we only have \$40,000 and the amount needed is about

**Stoichiometry Stumper #3 by Jadine Lee - Prezi**  
There are 3 astronauts in space that need to get rid of carbon dioxide. They emit 3,000 grams of carbon dioxide over the course of 3 days. They have 5,000 grams of sodium hydroxide which can be used to get rid of the carbon dioxide using this reaction: NaOH+CO2->Na2CO3+H2O. They have 2 days left before they land on earth.

**Stoichiometry Stumper #2 by Kailin Thomas**  
Can you explain the steps to solving these problems, I'm trying to study for a test but these three keep stumping me... The answers are 1. 359.9 g/mol 2. C4H7 3. 164g. Can you explain how to get those? 1. Cortisone consists of molecules each of which contains 21 atoms of carbon. The mass percentage of carbon in cortisone is 69.98% . What is the molecular mass of cortisone? 2.A 2.00-g sample of a ...

**Chemistry Stoichiometry Stumpers? | Yahoo Answers**  
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**Stoichiometry Stumper 1 Answer - electionsdev.calmatters.org**  
This was a problem in my chemistry review packet. It's the only one I don't know I don't know how to do. I've been working on it for hours. A mixture of methane (CH4) and ethane (C2H6) is burned. The two combustion reactions require a total of 179.959 grams of oxygen, and they produce a total of 226.473 grams of products. From this information, calculate the masses of methane and ethane in the ...

**Stoichiometry Stumper? | Yahoo Answers**  
Honors Stoichiometry 1 ... Underline or circle the final answer 1 Point 2 Point 3 Point. ... Honors Stoichiometry 10 Stoichiometry Stumper 1: Space Exploration In space shuttles, the CO 2 that the crew exhales is removed from the air by a reaction within canisters of ...

**Stoichiometry - Windsor Locks Public Schools**  
3 1 1.47 30.0 g 122.55KCl 1 mol KClO mol KCl 3 g KCl 1 3KCl mol KClOmol = g KClO49.3 74.55 2 2.73 g 1KClO 3 mol KCl molg KClO KCl 3 2 mol KCl O 3 = g KCl 122.55 1 1.66 74.55 g2 KCl KCl 42.7 g Fe mole Fe mol2 Fe 2O 3 g Fe Femol FeO 3 g Fe 2O 3 = g Fe 2O 3 55.85 4 1 61.0 17.0 g O 2 1 mol 2O 2 mol Fe 2O 3 g O32.00 3 mol O 2 g Fe 2O 3 g Fe 2O 3 = 1 ...

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