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The thermal energy of a monatomic gas is $\frac{3}{2} NkT = \frac{3}{2} PV$ so the change in thermal energy due to the expansion is $\Delta U = \frac{3}{2} P \Delta V = \frac{3}{2} P (V_f - V_i)$ (9)
 $= \frac{3}{2} (3.1 \times 10^5 \text{ Pa}) (3.1 \times 10^{-3} \text{ m}^3 - 1.0 \times 10^{-3} \text{ m}^3) = 1215.6 \text{ J}$ (11) The heat added to the gas must be enough to do the work plus increase the thermal energy, so $Q = \Delta U + W = 1620.8 \text{ J}$ (12) Example 2.

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Thermodynamic identity (Schroeder 3.4) $dU = TdS - pdV$ (1.5) Total differentials When we have a total differential, the things in front of the dS, dV , etc are partial derivatives. $T = \left(\frac{\partial U}{\partial S}\right)_V - p = \left(\frac{\partial U}{\partial V}\right)_S$ (1.6) Also, you can integrate along a path using a total differential, and can do linear algebra using total differential equations, e.g. substituting one for another.

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